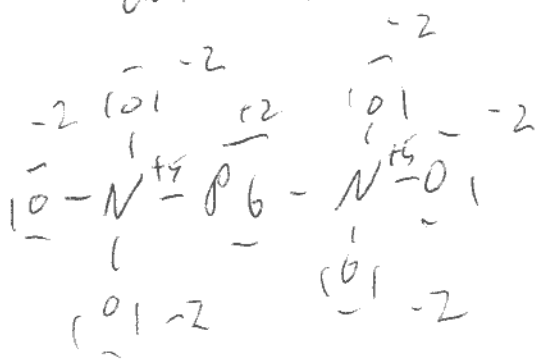
oxidiert

$$\overset{+2}{Pb}(\overset{+5}{N}\overset{-2}{O_3})_2 \rightarrow \overset{+4}{N}\overset{-2}{O_2} + \overset{+2}{Pb}\overset{-2}{O} + \overset{0}{O_2}$$

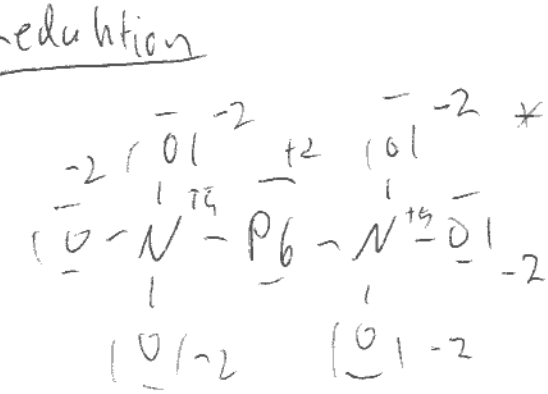
reduziert

Oxidation:



$$\rightarrow 2\,\overset{+4}{N}\overset{-2}{O_2} + \overset{+2}{Pb}\overset{-2}{O} + \langle O^{2-} \rangle + 2e^-$$

Reduktion



$$+ \langle O^{-2} \rangle + 2e^- \rightarrow O_2 + PbO + NO_2$$

* Dieser Reaktionsschritt liefert zusätzlich $2e^-$ daher sind nun $4e^-$ da um $2O^{2-}$ zu O_2 zu reduzieren.

$$2Pb(NO_3)_2 \rightarrow 4NO_2 + 2PbO + O_2$$